Chemistry 20	Unit 2
Lesson 1 - Introduction to Boyle's Law	84 mins

Gas Chemistry

What is pressure?

- Related to force
- Related to surface area
- Measured in Pascals (Pa)

What is atmospheric pressure?

- Pressure exerted by earth's atmosphere
- Greater altitude = Lower Pressure

How do we measure gas pressure?

- With a barometer.

What is the base Air Pressure?

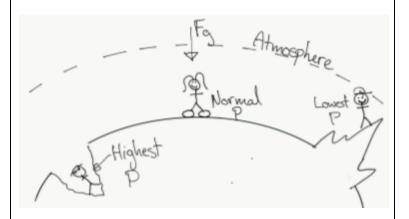
Sea Level

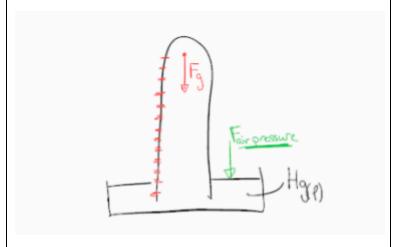
$$P = \frac{F}{4}$$

 $P = \frac{F}{A}$ F = Force(N)

$$A = Area(m^2)$$

$$P = Pressure\left(\frac{N}{m^2} = Pa\right)$$





If
$$F_0 = F$$

If $F_g = F_A$ - $Hg_{(l)}$ level is constant

If $F_g > F_A$ - $Hg_{(l)}$ level is drops

If $F_g < F_A$ - $Hg_{(l)}$ level is rises

If
$$F_g > F_A$$

$$f F_{\alpha} < F_{A}$$

1.000 atm (atmospheres)

760.00 mmHg (millimeters of mercury)

101.325 kPa (kilopascals)

Boyle's Law (Scuba Diver's Law)

- External P and V are <u>inversely</u> proportional

