| Chemistry 20 | Unit 2 |
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| Lesson 1 - Introduction to Boyle's Law | 84 mins |
| Gas Chemistry |  |
| What is pressure? <br> - Related to force <br> - Related to surface area <br> - Measured in Pascals (Pa) | $\begin{aligned} & P=\frac{F}{A} \\ & F=\operatorname{Force}(N) \\ & A=\operatorname{Area}\left(m^{2}\right) \\ & P=\operatorname{Pressure}\left(\frac{N}{m^{2}}=P a\right) \end{aligned}$ |
| What is atmospheric pressure? <br> - Pressure exerted by earth's atmosphere <br> - Greater altitude = Lower Pressure |  |
| How do we measure gas pressure? <br> - With a barometer. | $\text { If } F_{g}=F_{A}$ <br> - $\mathrm{Hg}_{(1)}$ level is constant $\text { If } F_{g}>F_{A}$ <br> - $\mathrm{Hg}_{(1)}$ level is drops $\text { If } F_{g}<F_{A}$ <br> - $\mathrm{Hg}_{(1)}$ level is rises |
| What is the base Air Pressure? <br> - Sea Level | 1.000 atm (atmospheres) <br> 760.00 mmHg (millimeters of mercury) <br> 101.325 kPa (kilopascals) |

Boyle's Law (Scuba Diver's Law)

- External P and V are inversely proportional

